

Binary Ionic Formula Practice

Name _____

Write the correct formula for each compound named below. Show the ions from which it is formed.

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|--|------------------|------------------|---------------|
| 1) sodium chloride | Na^{+1} | Cl^{-1} | NaCl |
| 2) lithium bromide | ----- | ----- | ----- |
| 3) magnesium flouride | ----- | ----- | ----- |
| 4) potassium oxide | ----- | ----- | ----- |
| 5) calcium sulfide | ----- | ----- | ----- |
| 6) aluminum iodide | ----- | ----- | ----- |
| 7) barium bromide | ----- | ----- | ----- |
| 8) aluminum sulfide | ----- | ----- | ----- |
| 9) calcium phosphide (P^{-3}) | ----- | ----- | ----- |
| 10) lithium selenide | ----- | ----- | ----- |
| 11) magnesium telluride | ----- | ----- | ----- |
| 12) aluminum fluoride | ----- | ----- | ----- |
| 13) lithium oxide | ----- | ----- | ----- |
| 14) beryllium iodide | ----- | ----- | ----- |

Binary Ionic Formula Practice

Name _____

1) sodium chloride	Na^{+1}	Cl^{-1}	NaCl
2) lithium bromide	Li^{+1}	Br^{-1}	LiBr
3) magnesium fluoride	Mg^{+2}	F^{-1}	MgF_2
4) potassium oxide	K^{+1}	O^{-2}	K_2O
5) calcium sulfide	Ca^{+2}	S^{-2}	CaS
6) aluminum iodide	Al^{+3}	I^{-1}	AlI_3
7) barium bromide	Ba^{+2}	Br^{-1}	BaBr_2
8) aluminum sulfide	Al^{+3}	S^{-2}	Al_2S_3
9) calcium phosphide (P^{-3})	Ca^{+2}	P^{-3}	Ca_3P_2
10) lithium selenide	Li^{+1}	Se^{-2}	Li_2Se
11) magnesium telluride	Mg^{+2}	Te^{-2}	MgTe
12) aluminum fluoride	Al^{+3}	F^{-1}	AlF_3
13) lithium oxide	Li^{+1}	O^{-2}	Li_2O
14) beryllium iodide	Be^{+2}	I^{-1}	BeI_2